PROCEEDINGS OF THE AMERICAN CHEMICAL SOCIETY.

Room 1, *University Building, N. Y.*

The last regular meeting was held Friday, October 6, 1882. The meeting was called to order at 8.45 P. M., Mr. P. Casamajor in the chair.

The minutes of the previous meeting were read, but as there was not a quorum present, they could not be acted upon.

After which, the first paper of the evening, "On the fractional dehydration of Ammonium Alum and the Atomicity of Aluminium," by Prof. J. W. Mallet, was read.

This paper describes a large number of interesting experiments, which go to show that Aluminium is either a triad or pentad.

After some remarks, the second paper "On the determination of Sulphur in Gas," by Mr. A. P. Hallock, was read.

Mr. Hallock stated that he used dehydrated sulphate of copper for the estimation of H_2S.

The apparatus used, consisted in a series of U tubes; the gas being allowed to pass through them for a certain length of time.

Dr. Grothe then inquired whether the above gas contained any Ammonia? Mr. Hallock stated that it did, but only in traces.

Dr. Grothe further asked whether the above gas contained any acetylene? Mr. Hallock replied that there was none. Upon which Mr. Casamajor remarked that possibly sheet rubber would remove small traces of hydrocarbons.

Mr. Elliott then stated that the Municipal Gas contains no other hydrocarbons than CH_4, and that this is not absorbed by the dehydrated sulphate of copper.

He further stated that all the sulphur in water gas is in the shape of sulphuretted hydrogen.

Some remarks about the combustion of sulphur in gas, lead to a lively discussion, in which Mr. Elliott stated that the H_2S was all burnt to H_2SO_4, while Dr. Grothe and Mr. Casamajor insisted that SO_2 is likewise formed.

The following gentlemen were then nominated:

Prof. Cuthbert P. Conrad, regular member. Proposed by J. W. Mallet, E. Waller and James H. Stebbins, Jr.

Ervin von Wilmouski, regular member. Proposed by James H. Stebbins, Jr., Wm. Rupp and Dr. Grothe.

Alfred L. Beebe, regular member. Proposed by J B. McIntosh, A. L. Colby and A. H. Elliott.

John Cawley, regular member. Proposed by James H. Stebbins, Jr., A. H. Elliott and P. Casamajor.

N. Hathaway, regular member. Proposed by E. Waller, Wm. Rupp and C. E. Munsell.

Geo. Leguin, associate member. Proposed by Nelson H. Darton, James H. Stebbins, Jr., and Wm. Rupp.

John G. Macfarlan, associate member. Proposed by P. Casamajor, James H. Stebbins, Jr., and A. H. Elliott.

Dr. Waller then stated on behalf of the committee on papers, that the lack of MS. produced much trouble, and he invited the members to assist said committee, by supplying as much MS. or abstracts as possible.

There being no further business the meeting adjourned.

JAMES H. STEBBINS, Jr.,
Recording Secretary.

Laboratory Notes.—By E. Waller.

PRECIPITATION OF BARIUM SULPHATE.

I have frequently observed that solutions of barium chloride seem to dissolve more air or carbon dioxide than most other solutions used as reagents, so much so that the addition of a cold barium solution to a boiling solution causes a brisk evolution of gas comparable with the effervescence due to the addition of acid to a carbonate. When a sulphate is present in the solution to which cold barium chloride solution has been added, even though the sulphate solution may have been boiling, the precipitate separates in a very finely divided condition, which gives much trouble in filtering, as the precipitate is so fine as to pass through the pores of a filter, and takes very long to settle.

If the barium solution is boiled for a few minutes before adding it, this phenomenon does not occur. The conclusion seems to be that the gases dissolved by the barium solution play an important part in the matter of the separation of the precipitate. Other chemists have evidently experienced the same trouble, for I have found recommendations in the literature to add a pinch of starch to the solution, and also more recently the addition of a little silver